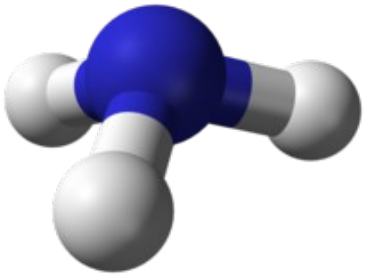
AMMONIA



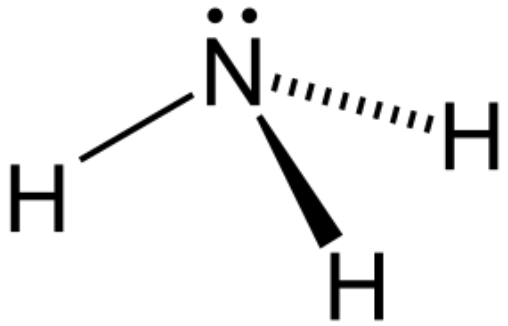
by Miles Mathis

First published January 31, 2024

I held this until February 2 so that I could do my biannual fund drive on Groundhog's Day, with a major new paper (two new papers, now) to salt it in. Just seemed like the thing to do. But I couldn't make the diagrams I wanted of Ammonia with Gimp and don't have Photoshop anymore, so I had to beg for *pro bono* help from a friend. Leading to some delay. Most of you know I only do this twice a year: post a simple reminder that donations are welcome. I don't do daily fund drives like most on the internet, don't do intrusive advertising or pop-ups, and don't bother you with merchandise, either. I think too much of you, and myself, to bother you with any of that. You can donate via Paypal, pay mileswmathis@yahoo.com, or mail me a check or other booty to POBox 335, Garden Valley CA 95633. Some imagine I can do this because I am privately wealthy—a trust-fund kid or something. Nope. I have been living hand-to-mouth for almost 40 years now, and although my paintings still give me a small income, the balance has been made up for about ten years by my readers like you. My science, art, and history readers help make this possible, allowing me to drive around the mainstream gatekeepers in several fields. I know that some of you have already donated in the past month or so, and I thank you for needing no reminder. The rest of you I thank in advance.

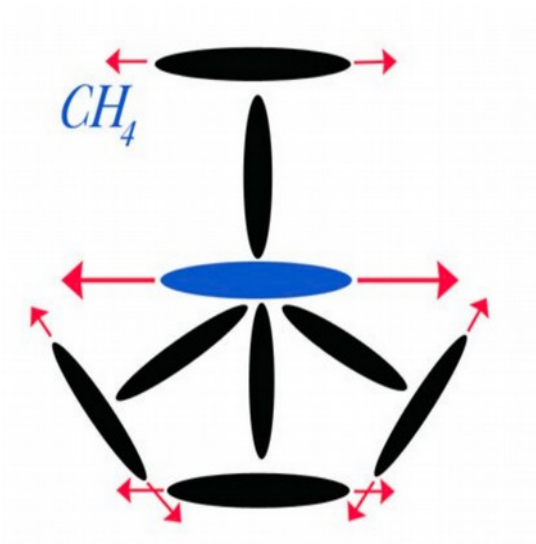
Many have asked me when I am going to get to the whole problem of acids and bases, since I have claimed electron bonding theory is a hoax from the ground up. If that is true, it implies all the historical explanations of acids are wrong, including those of Arrhenius, Bronsted, and most importantly Lewis. That's true, that exactly what that implies, and I will hit it now.

That diagram is the stick-and-ball diagram of ammonia, and it tells us at a glance. . . why the mainstream hasn't been able to solve this in almost 150 years. To solve it, we need a better diagram, one that includes the alphas and protons in the nuclei involved, as well as the charge channels. But the mainstream doesn't have that, so they have to explain everything with fake electron orbitals they made up from nothing almost as long ago as Arrhenius.



That's the other state-of-the-art mainstream diagram, which tells us the three bonds are different, but doesn't tell us why. Without a diagram of Nitrogen's nucleus, you can't possibly explain that, but since I have that diagram, I can. What I will do is follow closely my diagram and explanation of Methane, from that 2013 paper. 2013? Has it really been that long? Have I really let Ammonia sit for over a decade, while having this groundbreaking explanation in my pouch ready to go? Unfortunately, yes. I could have done this in 2013, but figured everyone with eyes open could already follow my procedure on their own. That or go rot.

Well, I haven't changed my mind on that: the mainstream can continue to go rot, which it is doing with ever greater speed. But for myself and my Muses I am back to get this in the can before I die.

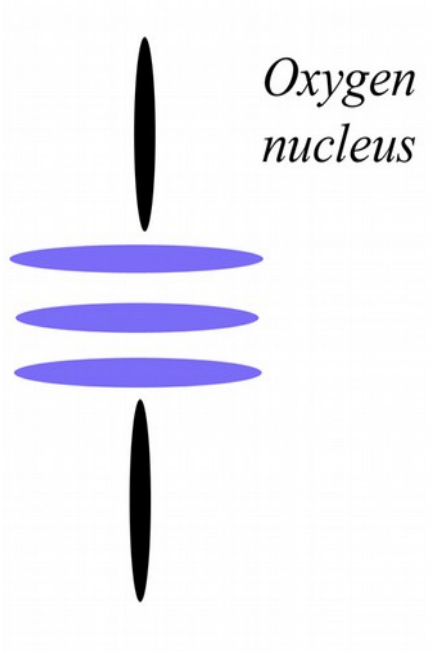


That is my final diagram of Methane from that paper, with the blue disk being the central alpha of Carbon and the black disks being protons. The Hydrogens are plugged into the existing polar charge channels of the Carbon nucleus. They do that because they are caught in those charge streams. It is all mechanical and physical, and everything happens here for a reason. Nothing is fit here to *ad hoc* structural or mathematical rules. The red arrows indicate the various charge channels out of the nucleus. The channels in are not drawn, but they come in from the poles, north and south.

You can see why my work on Methane helps us get into this, since Nitrogen is just one up from Carbon, and both Methane and Ammonia are built on the tetrahedron according to VSEPR. So let's begin by diagramming Nitrogen as three blue alphas in a stack, to make the core. They are linked and held together by charge moving pole to equator in both directions, through the alphas; and also by

charge moving pole-to-pole through the alphas, south to north and north to south. The south to north channel is stronger, but both exist, giving us both charge and anticharge.

That gives us six protons and six neutrons, leaving one each. We plug the seventh proton in the south pole and the neutron in the north, to create balance. So Nitrogen looks like Oxygen, but with a neutron in the north pole instead of a proton.



We plug the proton in the south to match the Earth, where the we find the south pole dominant. Since the alpha can accept two protons in the plug (roughly), we have room for one more proton in the south pole and two in the north. Since we are building a molecule, not a bigger nucleus, we aren't really plugging these three Hydrogens into those slots. They are in molecular positions, backed out a distance, but following those polar charge streams, which create vortices that reach well out into the ambient field around the nucleus. For this reason they do not create a charge stream as strong as the elemental one at the south pole. That seventh proton is in tight, so it channels with less loss than the backed-out Hydrogens in molecular positions.

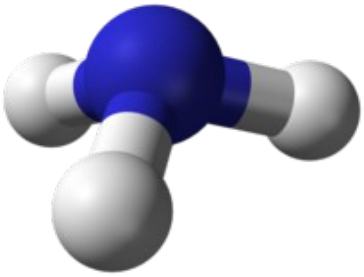
The south Hydrogen will align to the proton already there, but in the north we have that neutron to work around. You might think they would split the neutron, one going on each side of it, but that isn't how it works. They come in on the same northern vortex, so they come in on the same line and want to go on the same side of the neutron for that reason. Only once they come down almost into position and find one another there do they start splitting from eachother, due to their own charge fields. By that time they are on the same side of the neutron. This is what gives Ammonia its strange polar structure and unexpected empty position at the top of the tetrahedron.

The two north Hydrogens split, driven apart by their charge streams, and they also get pushed over away from the neutron's channel, which was initially moving straight up. That channel also gets pushed over in the other direction, but since it is the main nuclear channel of Nitrogen, it is stronger than that of the backed-out Hydrogens, even acting together.

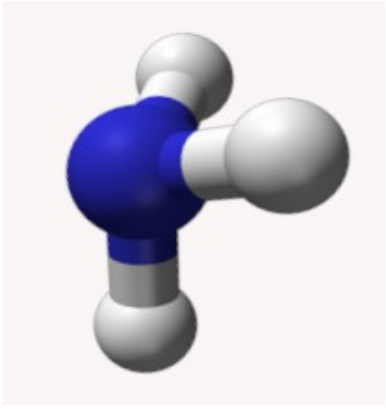
The neutron channel is releasing charge, so we ignore it in the diagram. We are only diagramming incoming charge streams, since obviously they create the bond. It is incoming charge that pulls the Hydrogens in. But notice already that two of our streams are technically anticharge, since they are coming in the north pole of the molecule. That will be important later.

If we diagram the neutron to the right, the two Hydrogen channels north are pushed over well to the left, 16.8 degrees from horizontal. What stops them there is not any VSEPR rules or tetrahedron composition rules. What stops them is the carousel level of Nitrogen, which, remember, is releasing equatorially, at 90 degrees all round in our 2D diagram. See the red arrows coming out of the blue alpha of Methane, above. That stream limits how much our northern Hydrogen streams can move away from the neutron, and also away from one another.

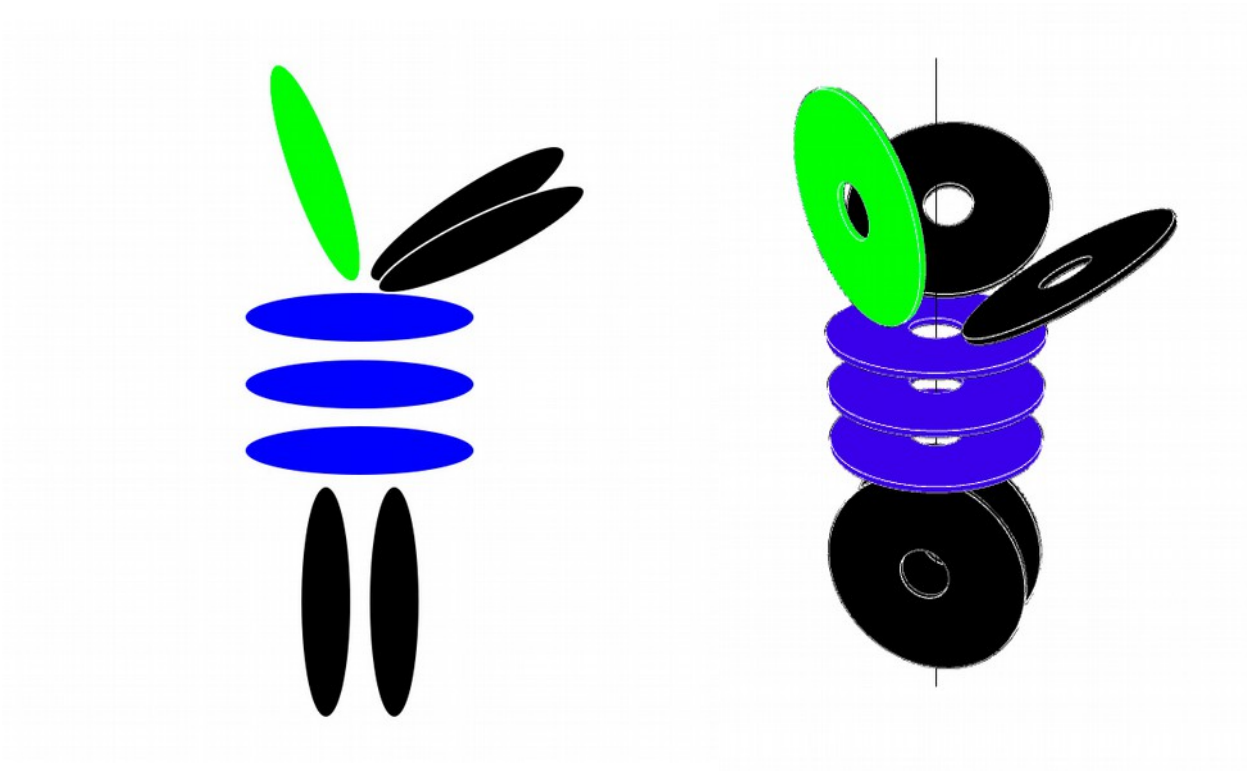
So what we find after all that is that the mainstream diagram is rotated, giving us a false impression of the molecule. We find this



should be this:



The way they have drawn it, it appears nothing is on the nuclear pole at the top, which is impossible. With my tweak, we now see the south channel/bond is primary, and why: it is on the nuclear pole of the Nitrogen, matching the charge stream of Nitrogen's seventh proton. And we can see that it is the neutron skewing the whole thing, since the unpictured neutron of Nitrogen is on the upper left side there, excluding the two upper Hydrogens. And that is why Ammonia's bond angle isn't 109.5: the neutron isn't an equivalent channeling or charge beast as the protons or Hydrogens. It won't act as an equal fourth leg of this tetrahedron, even if we draw it in.



That's my diagram of Ammonia, plus Arlo Emerson's fancy render in 3D. The green disk is a neutron, but I just drew it in for fun. It isn't doing much here. Once the protons plug in the north pole they overpower it and push it over.

Of course that means the mainstream diagram of NH4 can't be right either. NH4 doesn't magically drop that neutron, so there is no way it can be a perfect tetrahedron, as we are told.

You may not understand how we have room for a third proton in the north here to make NH4, but remember, Ammonia has two protons south: one of the Hydrogens and also the one belonging to Nitrogen. So although you can't see that in the mainstream diagram, Ammonia is still bottom heavy regarding charge balance. It has one strong proton south and one weak, while north there are two weak protons. So if we say the strong proton is 1 and the weak are .6, the charge strength south is 1.6 and north is 1.2. The Ammonia is a base, so it can take that other proton, taking the north up to 1.8. But then the north is too strong and Ammonium wants to release that proton to match the charge profile of the ambient field, which favors charge over anticharge. Which is why NH4 is an acid. That's how it really works. It has nothing to do with electron pairs.

Now that we know that, let's go back to the mainstream explanation. Let's see what they say about that 106.8 degree bond angle in Ammonia.

The ammonia molecule has a trigonal pyramidal shape, as predicted by the valence shell electron pair repulsion theory (VSEPR theory) with an experimentally determined bond angle of 106.7°. [29] The central nitrogen atom has five outer electrons with an additional electron from each hydrogen atom. This gives a total of eight electrons, or four electron pairs that are arranged tetrahedrally. Three of these electron pairs are used as bond pairs, which leaves one lone pair of electrons. The lone pair repels more strongly than bond pairs; therefore, the bond angle is

not 109.5°, as expected for a regular tetrahedral arrangement, but 106.8°.[29] This shape gives the molecule a dipole moment and makes it polar.

None of that makes any sense. The Nitrogen might have five outer electrons, if we take the one with the south proton and the two outer ones with the top and bottom alphas. But there is no reason Nitrogen would want the electrons of the three Hydrogens. To start with, those three electrons aren't loose. They should be bound to their protons strongly. There is no reason we have ever been given for Hydrogen to shed all its electrons in the presence of Nitrogen or anything else, and the reverse is also true: there is no reason for Nitrogen to shed its outer electrons so that they can be repaired to create the bonds. Not only is there no reason, we know experimentally it doesn't happen. Nitrogen can be monitored in the presence of Hydrogen, and no one has ever seen it shedding electrons like this. Beyond that, Nitrogen's electrons are already paired in the alphas, and it has no reason to attract any free electrons, from Hydrogen or elsewhere. And if it did (it doesn't), it would only want one to pair with its seventh electron.

It doesn't, because electrons don't pair up anyway. It goes against the charge definitions: electrons repel one another, they don't pair up and create bonds by magic. How do electrons pairing up create molecular bonds? Is it a matter of crossing paths, or what? You can't create bonds by pairing up like-charge particles. It is absurd. If you could, you could create the strong force by fiat as well: just pair up protons and say you are done.

Let me say it a different way. These bonds are all linear. Meaning, they create these bonds that we can draw a line through, calculating an angle, right? Each Hydrogen is in some position relative to the Nitrogen. This position is not a cloud or a probability, otherwise you couldn't calculate, much less measure, a bond angle. So how do electrons in squishy orbital clouds release upon bonding, creating pairs in some squishier meeting, and thereby create a real linear bond, of determinate size, strength, and position? With my charge channeling, this question is answered simply and directly: electrons have nothing to do with it. Charge is winds of real photons, and the nucleus channels them in real and defined paths. It is these pathways that are the bonds. But the mainstream has no answer to any of this.

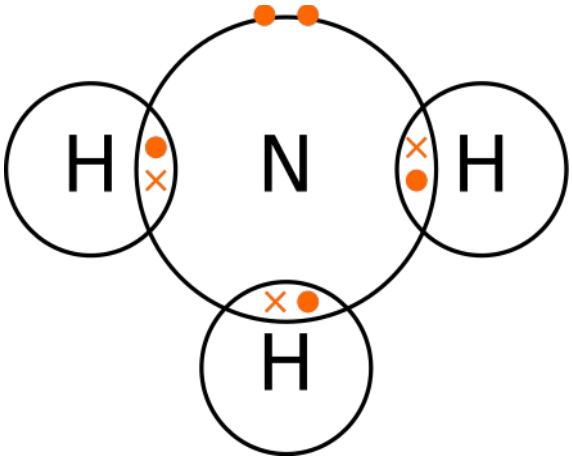
And it gets worse. We are told four electron pairs are created, but only three of them create bonds. The fourth pair doesn't have any bond to create, so we aren't sure why it was created in the first place. Electrons just like to pair up to suit theorists, I guess. They have to do something with it, so they use it to create the unexpected bond angle. It “repels more strongly than bond pairs”. What? Repels what more strongly? Protons, I guess, because that is the only way for it to affect bond angles. So in this novel theory, electrons attract other electrons, and repel protons.

Are you laughing yet? If not you aren't paying attention.

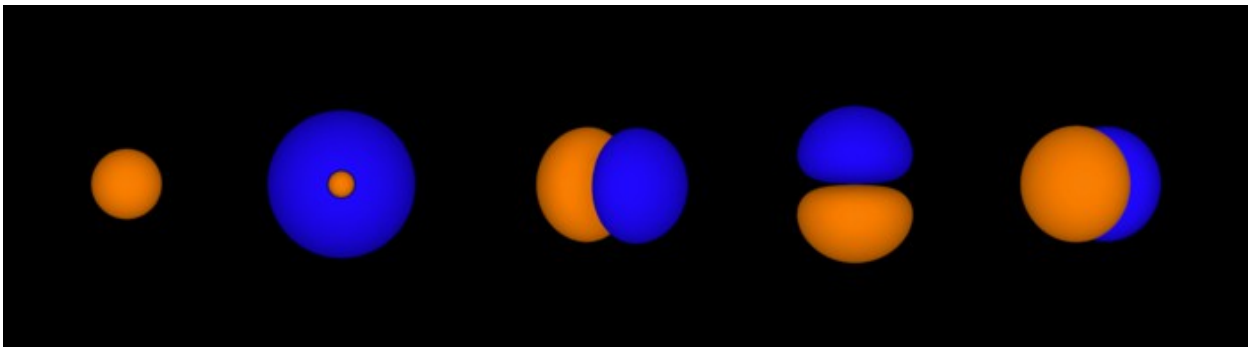
When electrons pair up for no reason, they apparently switch charge, suddenly repelling protons instead of being attracted to them. These people will say anything.

But ask yourself this: where does this fourth pair of electrons exist? The other three must exist in the bond lines, though we don't know how they got there or why. But the fourth pair, where does it exist? Well, if the bond angle changes in a tetrahedron, or even a partial tetrahedron, it must change in all spots equally, or we must be told why not. We have three bond angles here, right? So we need three changes from 109.5 in three different places. So this fourth electron pair must be in all three positions, changing that bond angle to 106.8.

The miracles just keep stacking up.



Here's another problem. They also give us that diagram of Ammonia, which shows the three unpaired electrons of Nitrogen in its *2p* orbitals bonding to the three Hydrogens. But its *2p* orbitals aren't arranged anything like that.

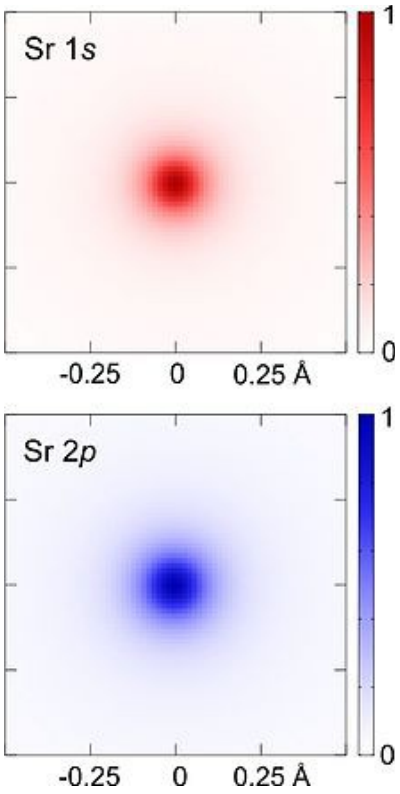


Those last three figures are the *2p* orbitals, x, y, and z. Or sort of front/back, top/bottom, left/right. Those are said to be the three valence or bonding electrons, not in actual orbits, but in standing barbell waves defined by quantum numbers, with *p* orbitals being $\ell = 1$. So how does the Hydrogen electron bond with any of that? Hydrogen's electron is in an *s* orbital, not a *p*, so how exactly do those two shapes bond? Their quantum numbers and standing waves are completely different, so it doesn't seem very likely they would pair up, much less bond. Pair up in what way, exactly? And bond in what way exactly? We have never been told.

So where did this theory of orbitals come from? It came from shapes seen in Hydrogen and alkali metals at different energy levels in spectroscopes in the early years of quantum theory. Early theorists simply guessed that those shapes were indication of electron orbitals, though they had no good reason to suppose that. When they needed a theory of molecular bonding, they manipulated those orbitals to manufacture a theory out of pretty much nothing, and have been adding to it ever since, building the greatest castle in the air in history. But I proved in 2011 there was no chance they were right, since those shapes they were seeing weren't electron orbitals, but shapes in the ambient field created around atoms by their charge fields. These shapes were indicating the shapes of the nuclei themselves, and the charge eddies at varying distances from the nucleus. Yes, free electrons could be influenced by those

shapes, but the electrons, free or bound, were not causing the shapes. As buoys in the field, they could not be causing anything like that, and therefore could also not be the cause of molecular bonds. As I showed, molecular bonds are created by charge channeling by the nucleus. The nucleus is like a pneumatic engine, with charge as the pneuma. Charge being strong winds of photons moving through and around all atoms.

Has the mainstream ever presented any proof or even indication atoms are bonding in this manner? No, the only indication we have ever been given is the bonds themselves. Molecules exist, therefore there must electron bonds. Even today, with powerful spectroscopy, we have no proof of electron bonding of any sort, and a lot of proof against it.

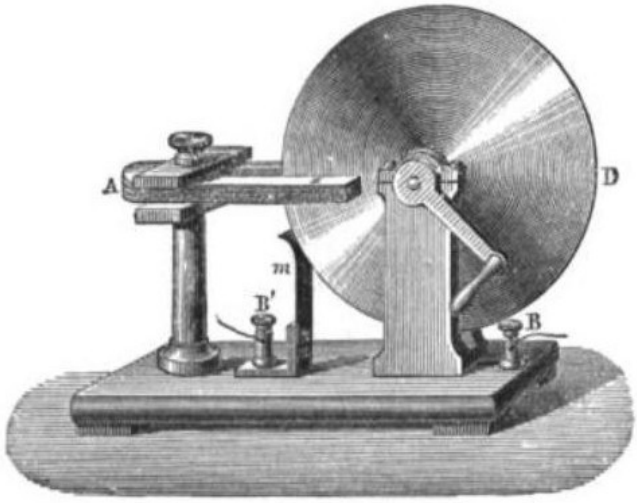


For instance, Wikipedia gives us those images from a recent EDX spectroscope of Strontium. Do you see anything but a red and blue dot? I don't. The blue dot is a little bigger, but otherwise those images don't tell us anything. They certainly aren't proof of electron orbitals or electron bonding theory.

As you now see, the mainstream theory of electron bonding fails on all levels, but it fails perhaps most spectacularly in explaining how atoms ever attract one another to start with. For electrons to pair up and bond requires very close proximity, so what provides the initial attraction, from a great distance, beyond the possible influence of bound electrons in charge balance with their nuclei? That question is especially raw with Hydrogen and Nitrogen, which, as gasses, have a low density. The spacing between atoms is great and the speeds are great as well, so how do we get the atoms together to start with? I guess the mainstream would have you believe it is all chance, and that the creation of Ammonia relies on random close passes, where electrons suddenly reach out and grab eachother. But in my theory, all that is explained very simply, since charge is everywhere. The charge streams of elements are very powerful and their vortices reach far out into the ambient field. In this way, larger

nuclei can easily pull in smaller ones, and Hydrogen, being a single proton, is especially easy to explain. It is pulled in by the much larger polar streams of Nitrogen, driven by those three stacked alphas in line. Besides, in the real world, we don't get Ammonia by combining gasses like that. They take something like Calcium Cyanamide and divide it with water to get Ammonia. CaCN2 is a salt, therefore a solid, and of course water is a liquid, so you aren't trying to bond gasses. And the Calcium adds to the charge field strength, helping draw the Hydrogen off water. Calcium overpowers Oxygen's charge field, allowing Hydrogen to move to Nitrogen.

It helps to think of charge as a particulate wind made of photons, and to think of alphas and protons as fans. They are fans set to what we would call reverse, drawing other quantum in and creating charge channels. The nucleus is a charge recycler. The fan analogy is useful, since although it is just a visualization, the channels really are created by spins, as with fans. And has long been known that spinning disks, even at the macrolevel, act like fans regarding charge and EM.



That's the old Faraday disk, which you can look up if you don't believe me.